

# Metallurgy Volume 1



This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1911 Excerpt: ...and the roasted ore, if the heat be not kept up artificially by the burning of outside fuel, cools off gradually, and appears quite black, to assume later a red color, natural to the ferric oxide into which it has become metamorphosed. The latter substance begins to lose oxygen at 1375 (Hofman) in the air, becoming magnetic. It sinters into a mass at 1300. The product of such a roasting consists of the oxide of iron, with a smaller quantity of basic sulphate, the trioxide amounting, often, to but 0.4 or 0.5 per cent. of the whole, corresponding to from 1/3 to 1/6 per cent. of ferrous sulphate. If the mass be again raised to a glowing heat, by the burning of outside fuel, it will increase in bulk and again become woolly, and sulphur trioxide will be driven away, leaving the iron with which it was combined in the form of ferric oxide, through the absorption of oxygen. If soot or reducing gases come in contact with the ferric oxide, they will change it to the magnetic form in part, by removing oxygen. It is understood that three different oxides of iron may be formed by such treatment of pyrite, namely, ferrous, ferric, and magnetic oxide. Regarding the former, Wedding asserted that while it was detectable in roasted material, it is not practicable to so conduct the operation as to produce it commercially. In practice the operation proceeds directly to the production of magnetic oxide, and without finding any particular stopping place, mounts to the ferric form, and all the iron reaches this stage if the roasting be prolonged sufficiently. By more exact experimentation different degrees of the magnetic oxidation may be distinguished, varying in the proportion of oxygen from

that containing six equivalents of iron to seven of oxygen to seven of iron to eight of...

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